

**APSET**

**Previous Year Paper**

**2014 Paper II  
Chemical Science**

|                           |                          |                                |
|---------------------------|--------------------------|--------------------------------|
| <b>SUBJECT CODE</b>       | <b>SUBJECT</b>           | <b>PAPER</b>                   |
| <b>A-02-02</b>            | <b>CHEMICAL SCIENCES</b> | <b>II</b>                      |
| <b>HALL TICKET NUMBER</b> |                          | <b>QUESTION BOOKLET NUMBER</b> |
|                           |                          | <b>501067</b>                  |
| <b>OMR SHEET NUMBER</b>   |                          |                                |
|                           |                          |                                |
| <b>DURATION</b>           | <b>MAXIMUM MARKS</b>     | <b>NUMBER OF PAGES</b>         |
| <b>1 HOUR 15 MINUTES</b>  | <b>100</b>               | <b>16</b>                      |
|                           |                          | <b>NUMBER OF QUESTIONS</b>     |
|                           |                          | <b>50</b>                      |

This is to certify that, the entries made in the above portion are correctly written and verified.

**Candidate's Signature**

**Name and Signature of Invigilator**

**Instructions for the Candidates**

- Write your Hall Ticket Number in the space provided on the top of this page.
- This paper consists of fifty multiple-choice type of questions.
- At the commencement of examination, the question booklet will be given to you. In the first 5 minutes, you are requested to **open the booklet and compulsorily examine it as below** :
  - To have access to the Question Booklet, tear off the paper seal on the edge of this cover page. Do not accept a booklet without sticker-seal and do not accept an open booklet.
  - Tally the number of pages and number of questions in the booklet with the information printed on the cover page. Faulty booklets due to pages/questions missing or duplicate or not in serial order or any other discrepancy should be got replaced immediately by a correct booklet from the invigilator within the period of 5 minutes. Afterwards, neither the Question Booklet will be replaced nor any extra time will be given.
  - After this verification is over, the Test Booklet Number should be entered in the OMR Sheet and the OMR Sheet Number should be entered on this Test Booklet.
- Each item has four alternative responses marked (A), (B), (C) and (D). You have to darken the circle as indicated below on the correct response against each item.  
**Example :** (A) (B) (C) (D)  
 where (C) is the correct response.
- Your responses to the items are to be indicated in the **OMR Answer Sheet given to you**. If you mark at any place other than in the circle in the Answer Sheet, it will not be evaluated.
- Read instructions given inside carefully.
- Rough Work is to be done in the end of this booklet.
- If you write your name or put any mark on any part of the OMR Answer Sheet, except for the space allotted for the relevant entries, which may disclose your identity, you will render yourself liable to disqualification.
- The candidate must handover the OMR Answer Sheet to the invigilators at the end of the examination compulsorily and must not carry it with you outside the Examination Hall. The candidate is allowed to take away the carbon copy of OMR Sheet and used Question paper booklet at the end of the examination.
- Use only Blue/Black Ball point pen.
- Use of any calculator or log table etc., is prohibited.
- There is no negative marks for incorrect answers.

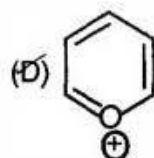
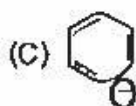
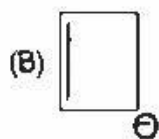
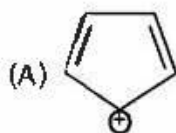
**అధ్యక్షులకు సూచనలు**

- ఈ పుట పై భాగంలో ఇవ్వబడిన స్థలంలో మీ హాల్ టికెట్ నంబరు రాయండి.
- ఈ ప్రశ్న పత్రము యాభై బహుళైచ్ఛిక ప్రశ్నలను కలిగి ఉంది.
- పరీక్ష ప్రారంభమున ఈ ప్రశ్నపత్రము మీకు ఇవ్వబడుతుంది. మొదటి ఐదు నిమిషములలో ఈ ప్రశ్నపత్రమును తెరిచి కింద తెలిపిన అంశాలను తప్పనిసరిగా పరిచూసుకోండి.
  - ఈ ప్రశ్న పత్రమును చూడడానికి కవర్ పేజీ అంచున ఉన్న కాగితపు సీలును చించండి. స్టిక్కర్ సీలులేని మరియు ఇవివరకే తెరిచి ఉన్న ప్రశ్నపత్రమును మీరు అంగీకరించవద్దు.
  - కవర్ పేజీ పై ముద్రించిన సమాచారం ప్రకారం ఈ ప్రశ్నపత్రములోని పేజీల సంఖ్యను మరియు ప్రశ్నల సంఖ్యను సరిచూసుకోండి. పేజీల సంఖ్యకు నెలబంధించి గానీ లేదా సూచించిన సంఖ్యలో ప్రశ్నలు లేకపోవుట లేదా నిజప్రతి కాకపోవుట లేదా ప్రశ్నలు క్రమవర్ధంతో లేకపోవుట లేదా ఏదైనా తేడాలుండటం వంటి దోషపూరితమైన ప్రశ్న పత్రాన్ని వెంటనే మొదటి ఐదు నిమిషాల్లో పరీక్షా పర్యవేక్షకునికి తిరిగి ఇచ్చివేసి దానికి బదులుగా సరిగ్గా ఉన్న ప్రశ్నపత్రాన్ని తీసుకోండి. తరువాతరం ప్రశ్నపత్రము మార్చబడదు అదనపు సమయం ఇవ్వబడదు.
  - పై విధంగా సరిచూసుకొన్న తర్వాత ప్రశ్నపత్రం సంఖ్యను OMR పత్రము పై అందించిన OMR పత్రము సంఖ్యను ఈ ప్రశ్నపత్రము పై నిర్ణయస్థలంలో రాయవలెను.
- ప్రతి ప్రశ్నకు నాలుగు ప్రతిస్పందనలు (A), (B), (C) మరియు (D) లుగా ఇవ్వబడ్డాయి. ప్రతిప్రశ్నకు నచ్చిన ప్రతిస్పందనను ఎన్నుకొని కింద తెలిపిన విధంగా OMR పత్రములో ప్రతి ప్రశ్నా సంఖ్యకు ఇవ్వబడిన నాలుగు వృత్తాల్లో నచ్చిన ప్రతిస్పందనను సూచించే వృత్తాన్ని బాల్ పాయింట్ పెన్ తో కింద తెలిపిన విధంగా పూరించాలి.  
**ఉదాహరణ :** (A) (B) (C) (D)  
 (C) నచ్చిన ప్రతిస్పందన అయితే
- ప్రశ్నలకు ప్రతిస్పందనలను ఈ ప్రశ్నపత్రములో ఇవ్వబడిన OMR పత్రము పైన ఇవ్వబడిన వృత్తాల్లోనే పూరించి గుర్తించాలి. అలాకాక సమాధాన పత్రంపై వేరొక చోట గుర్తిస్తే మీ ప్రతిస్పందన మూల్యాంకనం చేయబడదు.
- ప్రశ్న పత్రము లోనం ఇచ్చిన సూచనలను జాగ్రత్తగా చదవండి.
- చిత్తుపనిని ప్రశ్నపత్రము చివర ఇచ్చిన ఖాళీస్థలములో చేయాలి.
- OMR పత్రము పై నిర్దేశ స్థలంలో సూచించవలసిన వివరాలు తప్పించి ఇతర స్థలంలో మీ గుర్తింపును తెలిపే విధంగా మీ పేరు రాయడం గానీ లేదా ఇతర చిహ్నాలను పెట్టడం గానీ చేసినట్లయితే మీ అనర్హతకు మీరే బాధ్యులవుతారు.
- పరీక్ష పూర్తయిన తర్వాత మీ OMR పత్రాన్ని తప్పనిసరిగా పరీక్ష పర్యవేక్షకుడికి ఇవ్వాలి. వాటిని పరీక్షగది బయటకు తీసుకువెళ్ళకూడదు. పరీక్ష పూర్తయిన తరువాత అధ్యక్షులు ప్రశ్న పత్రాన్ని OMR పత్రం యొక్క కార్బన్ కాపీ తీసుకువెళ్ళవచ్చు.
- నీలి/నల్ల రంగు బాల్ పాయింట్ పెన్ మాత్రమే ఉపయోగించాలి.
- దానిధర్మమే లేకున్న, క్వాలిటీకంట్రోలు, ఎంక్లౌజర్డ్ పరికరాలు మొదలగునవి పరీక్షగదిలో ఉపయోగించడం నిషేధం.
- తప్పని సమాధానాలకు మార్కుల తగ్గింపు లేదు.

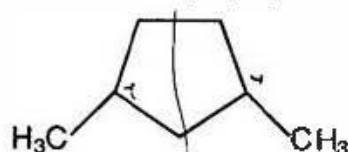


CHEMICAL SCIENCE  
Paper – II

1. Identify the aromatic compound.



2. How many stereoisomers are possible for 1,3-dimethylcyclopentane?



- (A) One  
(B) Two  
(C) Three  
(D) Four

3. The correct order of the solubility of the following in water is

- (A)  $\text{BeSO}_4 < \text{MgSO}_4 < \text{CaSO}_4 < \text{SrSO}_4$   
(B)  $\text{BeSO}_4 > \text{MgSO}_4 > \text{CaSO}_4 > \text{SrSO}_4$   
(C)  $\text{BeSO}_4 < \text{MgSO}_4 = \text{CaSO}_4 < \text{SrSO}_4$   
(D)  $\text{BeSO}_4 > \text{MgSO}_4 = \text{CaSO}_4 < \text{SrSO}_4$

4. Match the following :

- |                            |                |
|----------------------------|----------------|
| I. Beckmann rearrangement  | 1. Oxetane     |
| II. Paterno-Büchi reaction | 2. Olefin      |
| III. Wittig reaction       | 3. Enamine     |
| IV. Stork reaction         | 4. Caprolactam |

- |     | I | II | III | IV |
|-----|---|----|-----|----|
| (A) | 1 | 2  | 3   | 4  |
| (B) | 2 | 3  | 4   | 1  |
| (C) | 4 | 1  | 2   | 3  |
| (D) | 3 | 2  | 1   | 4  |

5. In a radical polymerization reaction,  $10^4$  monomer molecules are consumed to form 10 polymer molecules in dt time. What is the degree of polymerization (DP) ?

- (A)  $10^5$   
(B)  $10^4$   
(C)  $10^3$   
(D)  $10^2$

6. The correct relationship between vibrational frequency ( $\bar{\nu}$ ) and force constant (k) is

(A)  $(\bar{\nu})^2 \propto k^2$

(B)  $(\bar{\nu}) \propto \frac{1}{k}$

(C)  $(\bar{\nu})^2 \propto k$

(D)  $(\bar{\nu}) \propto (k)^{-\frac{1}{2}}$

7. Gibbs-Helmholtz equation is expressed as

(A)  $\left[ \frac{\partial(\Delta G/T)}{\partial T} \right]_P = \frac{-\Delta H}{T^2}$

(B)  $\left[ \frac{\partial(\Delta G/T)}{\partial T} \right]_V = \frac{-\Delta E}{T^2}$

(C)  $dG = -SdT + vdP$

(D)  $dH = Tds + vdP$

8. The number of radial nodes for a 3s orbital is

(A) 0

(B) 1

(C) 2

(D) 3

9. Wave function of a particle at some point x has the value  $\psi$ , then the probability of finding the particle between x and x+dx is proportional to

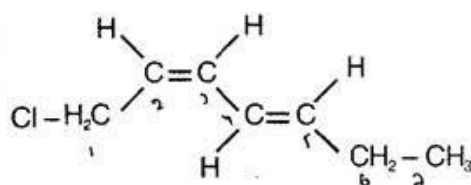
(A)  $|\psi| dx$

(B)  $|\psi|^3 dx$

(C)  $|\psi|^2 dx$

(D)  $|\psi|$

10. Give the IUPAC name for the given structure



(A) 1-Chloro - 2Z,4Z - heptadiene

(B) 1-Chloro - 2Z,4E - 2,4 - heptadiene

(C) 7-Chloro - 2Z,4E - heptadiene

(D) 1-Chloro - 2E,4Z - 2,4 - heptadiene

11. Which of the following amino acid belongs to neutral category ?

I. Glycine

II. Arginine

III. Tyrosine

IV. Asparagine

(A) I, II

(B) II, III

(C) III, IV

(D) I, III



12. The correct decreasing order of the first ionization energy of the elements P, S, Cl is

- (A) Cl, S, P
- (B) Cl, P, S
- (C) P, Cl, S
- (D) S, Cl, P

13. Which of the following does not obey 18-electron rule?

- (A)  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$
- (B)  $\text{Fe}(\text{CO})_5$
- (C)  $\text{Mn}(\text{CO})_5$
- (D)  $\text{Ru}(\text{CO})_5$

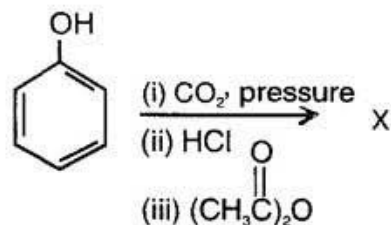
14. Zero point energy ( $E_0$ ) of a harmonic oscillator is

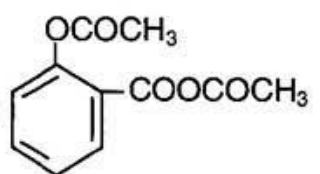
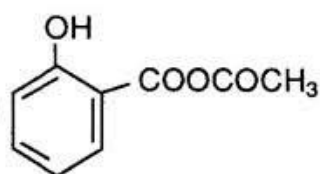
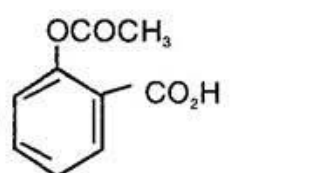
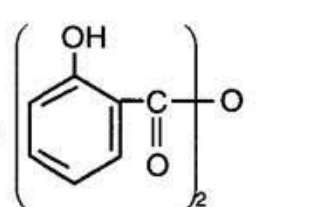
- (A)  $E_0 = h^2/8 mL^2$
- (B)  $E_0 = \frac{1}{2} h \nu$
- (C)  $E_0 = 0$
- (D)  $E_0 = 4h^2/8 mL^2$

15. The ground state term symbol of Li is

- (A)  $^2S_{\frac{1}{2}}$
- (B)  $^2P_{\frac{1}{2}}$
- (C)  $^1S_{\frac{1}{2}}$
- (D)  $^1D_2$

16. The major product in the following reaction is



- (A) 
- (B) 
- (C) 
- (D) 

17. The catalyst produced by treating  $\text{TiCl}_4$  with  $(\text{C}_2\text{H}_5)_3\text{Al}$  is

- (A) Wilkinson's catalyst
- (B) Ziegler – Natta catalyst
- (C) ZSM – 5 catalyst
- (D)  $[\text{PdCl}_4]^{2-}$  catalyst

18. Electronic configuration of  $\text{Gd}^{+3}$  is

- (A)  $[\text{Rn}] 4f^7$
- (B)  $[\text{Rn}] 4f^{10}$
- (C)  $[\text{Xe}] 4f^7$
- (D)  $[\text{Xe}] 4f^{10}$

19. The most stable conformer is

- (A)
- (B)
- (C)
- (D)

20. The isoelectronic compound with naphthalene is

- (A)
- (B)
- (C)
- (D)

21. According to MO theory, the strength of bonding in  $\text{O}_2^+$ ,  $\text{O}_2$  and  $\text{O}_2^-$  in follows the order:

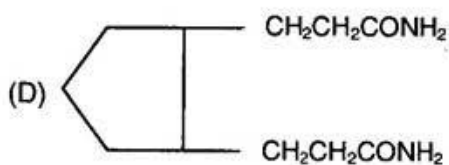
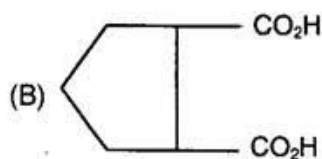
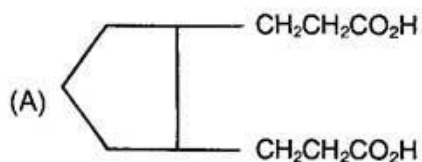
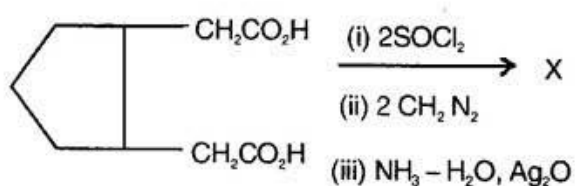
- (A)  $\text{O}_2^+ > \text{O}_2 > \text{O}_2^-$
- (B)  $\text{O}_2^- > \text{O}_2 > \text{O}_2^+$
- (C)  $\text{O}_2 > \text{O}_2^+ > \text{O}_2^-$
- (D)  $\text{O}_2 > \text{O}_2^- > \text{O}_2^+$



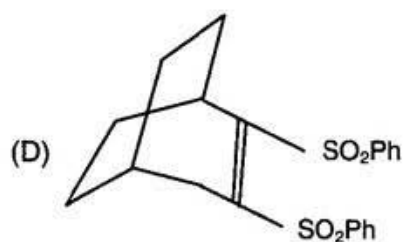
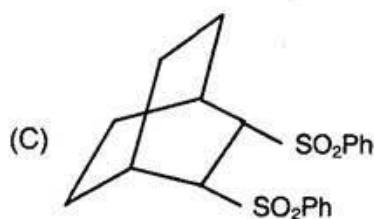
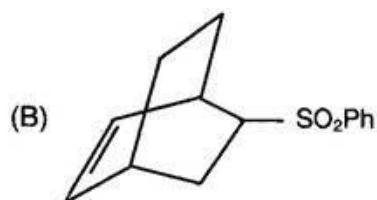
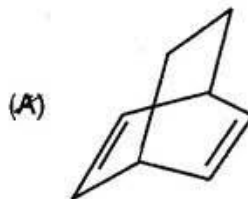
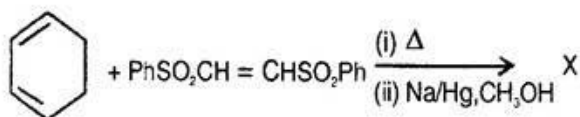
22. What is the point group of trans 1, 2-dichloro ethene ?

- (A)  $C_{2v}$  (B)  $C_{2h}$   
(C)  $D_{2h}$  (D)  $D_{2d}$

23. The major product in the following reactions is



24. The major product in the following reaction is



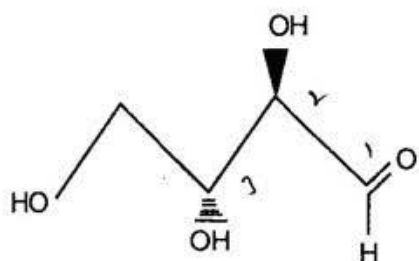
25. Number of B – B bonds present in tetraborane,  $B_4H_{10}$  is

- (A) 1
- (B) 2
- (C) 3
- (D) 4

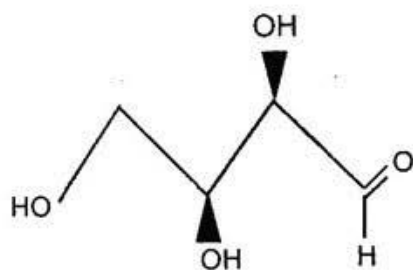
26. The metal atom present in Wilkinson's catalyst is

- (A) W
- (B) Rh
- (C) Pt
- (D) Ir

27. The following two compounds are



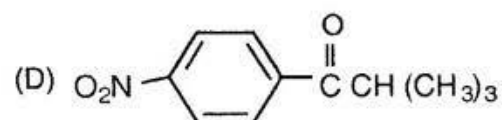
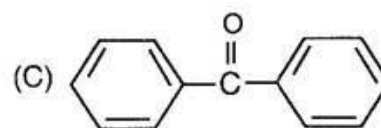
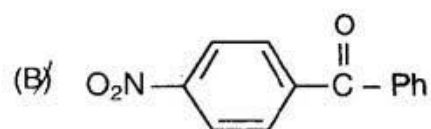
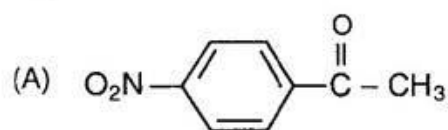
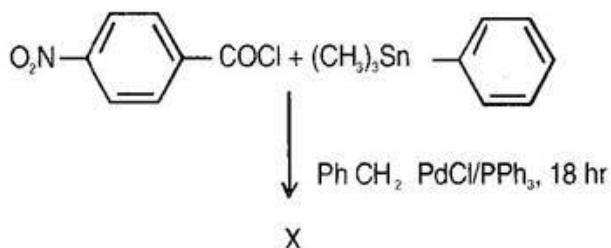
and



- (A) C(3) epimers
- (B) C(2) epimers
- (C) Different configuration at C(2)
- (D) Mesomers

28. The major product in the following

reaction is



29. The gas released by PS-II in photosynthesis process that occurs in green plants is

- (A)  $O_2$
- (B)  $N_2$
- (C) CO
- (D)  $CO_2$



30. For a  $\text{Co}^{3+}$  octahedral complex, the mean pairing energy is  $17400 \text{ cm}^{-1}$  and  $\Delta_0$  is  $24700 \text{ cm}^{-1}$ . The CFSE of the complex in Dq units is

- (A) 0 Dq  
(B) -6 Dq  
(C) -24 Dq + 2P  
(D) -12 Dq + 3p

31. The number of degrees of freedom at water

(I)  $\rightleftharpoons$  water (g) equilibrium is

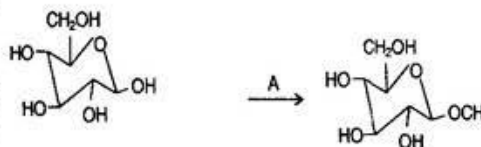
- (A) 0  
(B) 1  
(C) 2  
(D) 3

32. The molar partition function (Q) for system of distinguishable independent molecules is related to the molecular partition (q) as

- (A)  $Q = q^N$   
(B)  $Q = Nq$   
(C)  $Q = RT \left( \frac{\partial \ln q}{\partial \ln v} \right)$   
(D)  $Q = \frac{q^N}{N}$



33. The reagent 'A' in the following reaction is



- (A)  $\text{CH}_3\text{I} / \text{Ag}_2\text{O}$   
(B)  $\text{CH}_3 - \overset{\text{O}}{\parallel} \text{C} - \text{O} - \overset{\text{O}}{\parallel} \text{C} - \text{CH}_3 / \text{pyridine}$   
(C)  $\text{CH}_3\text{OH} / \text{H}^+$   
(D)  $\text{CH}_3 - \overset{\text{O}}{\parallel} \text{C} - \text{O} - \overset{\text{O}}{\parallel} \text{C} - \text{CH}_3 / \text{OH}^-$

34. Which of the following reactions is an atom economy reaction ?

- (A) Claisen condensation  
(B) Friedel Crafts reaction  
(C) Diels Alder reaction  
(D) Perkin reaction

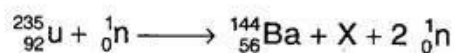
35. Which of the following solvents is an environmentally benign ?

- (A) Acetone  
(B) Acetic acid  
(C) Water  
(D) Tetrahydro furan

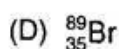
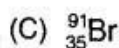
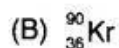
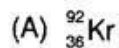




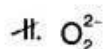
36. In the given nuclear fission reaction



identify X from the following



37. Which among the following has a bond order equal to 1 ?



The correct combination is

(A) I, II

(B) I, IV

(C) II, IV

(D) III, IV

38. An alkaloid containing

N-methylpiperidine moiety exhibits

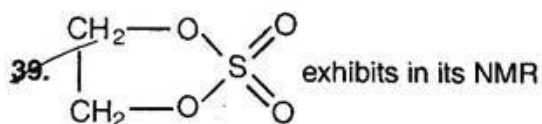
Narcotic analgesic activity is

(A) Atropine

(B) Nicotine

(C) Morphine

(D) Quinine



(A) Two singlets

(B) Two doublets

(C) Two triplets

(D) One singlet

40. Camphor belongs to

(A) Acyclic monoterpene

(B) Monocyclic monoterpene

(C) Bicyclic monoterpene

(D) Acyclic sesquiterpene



41. The number of ESR lines obtained for hydrogen atom due to hyperfine coupling is

- (A) 1
- (B) 2
- (C) 3
- (D) 4

42. Which of the following is a hard acid-soft base pair ?

- (A)  $\text{La}^{+3}$  and  $\text{OH}^-$
- (B)  $\text{Co}^{+2}$  and  $\text{SO}_3^{-2}$
- (C)  $\text{Pt}^{+2}$  and  $\text{PO}_4^{3-}$
- (D)  $\text{La}^{+3}$  and  $\text{CN}^-$

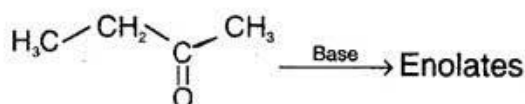
43. According to Debye-Huckel theory of strong electrolytes the electrophoretic effect which reduces the velocity of the ions is due to

- (A) Interionic attraction
- (B) Solvation of ions
- (C) Ion pair formation
- (D) Viscosity of the solvent

44. The rate constant of a diffusion controlled reaction ( $k_D$ ) is related to the viscosity of the medium ( $\eta$ ) in the following manner

- (A)  $K_D \propto \eta$
- (B)  $K_D \propto \eta^{-1}$
- (C)  $K_D \propto \eta^2$
- (D)  $K_D \propto \eta^{-2}$

45. How many types of enolates are possible by treating ethylmethylketone with a base ?



- (A) One
- (B) Two
- (C) Three
- (D) Four

46. Indicate the types of pericyclic reactions in the following conversions.



- (A) Electrocyclic reactions
- (B) Sigmatropic shifts
- (C) Electrocyclic and sigmatropic shift
- (D) Cycloadditions



47. Strongest super acid among the following is

- (A)  $\text{SbF}_5$  in HI
- (B)  $\text{SbF}_5$  in HF
- (C)  $\text{SbCl}_5$  in HI
- (D)  $\text{SbCl}_5$  in HCl

48. The coordination number of iron in cytochrome 'C' is

- (A) 4
- (B) 5
- (C) 6
- (D) 3

49. The correct statement regarding adsorption of gases on solid surfaces is

- (A) Langmuir and BET equations are applicable to monolayer adsorption
- (B) Langmuir and BET equations are applicable to multilayer adsorption
- (C) Langmuir equation is applicable to monolayer adsorption and BET equation is applicable to multilayer adsorption
- (D) Langmuir equation is applicable to multilayer adsorption and BET equation is applicable to monolayer adsorption

50. According to band theory of solids, the donor band in a n-type semiconductor is

- (A) close to valence band of parent material
- (B) close to conduction band of parent material
- (C) mixed with valence band of parent material
- (D) mixed with conduction band of parent material